**Figure: 30 TAC §112.203(a)(2)**

$$Dec\%_{sulfolene} = \frac{8.821921}{[(1 + e^{12.429479+(0.007527 \times time)}) \times (1 + e^{8.743395-(0.071099 \times temp)})]}$$

Where

$Dec\%_{sulfolene}$ = the percentage of the weight of sulfolene that decomposes;

$e$ = Euler's number, which is a mathematical constant approximately equal to 2.71828;

$time$ = the number of hours that the sulfolene has been at the monitored temperature; and

$temp$ = the monitored temperature in degrees Fahrenheit.